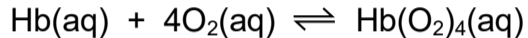


- 12** One molecule of haemoglobin, Hb, can bind with four molecules of oxygen according to the equation shown.



When the equilibrium concentration of O_2 is $7.6 \times 10^{-6} \text{ mol dm}^{-3}$, the equilibrium concentrations of Hb and $\text{Hb(O}_2)_4$ are equal.

What is the numerical value of K_c for this equilibrium?

- A** 3.0×10^{20} **B** 1.3×10^5 **C** 7.6×10^{-6} **D** 3.3×10^{-21}