

CHEMISTRY

Section II

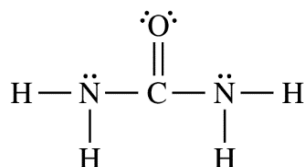
Time—1 hour and 45 minutes

7 Questions

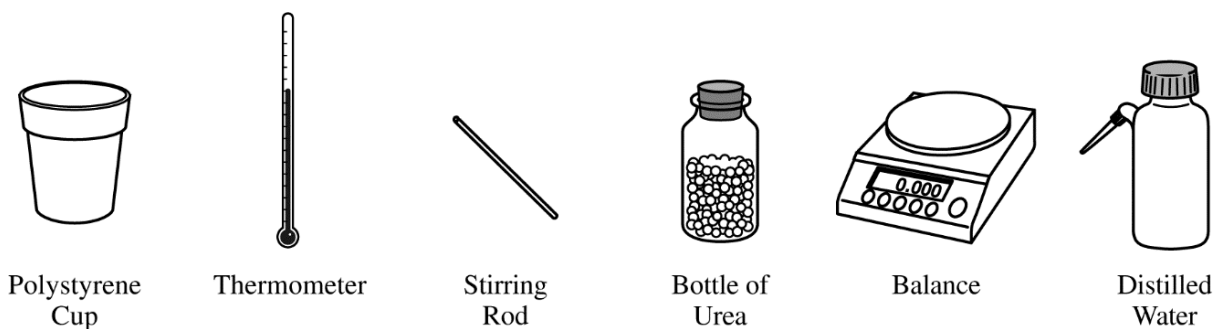
YOU MAY USE YOUR CALCULATOR FOR THIS SECTION.

Directions: Questions 1–3 are long free-response questions that require about 23 minutes each to answer and are worth 10 points each. Questions 4–7 are short free-response questions that require about 9 minutes each to answer and are worth 4 points each.

Write your response in the space provided following each question. Examples and equations may be included in your responses where appropriate. For calculations, clearly show the method used and the steps involved in arriving at your answers. You must show your work to receive credit for your answer. Pay attention to significant figures.



1. The compound urea, H_2NCONH_2 , is widely used in chemical fertilizers. The complete Lewis electron-dot diagram for the urea molecule is shown above.
 - (a) Identify the hybridization of the valence orbitals of the carbon atom in the urea molecule.
 - (c) Calculate the concentration of urea, in mol/L, in the saturated solution at $20.^\circ\text{C}$.
 - (d) The student also determines that the concentration of urea in a saturated solution at 25°C is 19.8 M . Based on this information, is the dissolution of urea endothermic or exothermic? Justify your answer in terms of Le Chatelier's principle.



- (e) The equipment shown above is provided so that the student can determine the value of the molar heat of solution for urea. Knowing that the specific heat of the solution is $4.18\text{ J/(g}\cdot^\circ\text{C)}$, list the specific measurements that are required to be made during the experiment.

	$S^\circ\text{ (J/(mol}\cdot\text{K))}$
$\text{H}_2\text{NCONH}_2(s)$	104.6
$\text{H}_2\text{NCONH}_2(aq)$	?

- (f) The entropy change for the dissolution of urea, $\Delta S^\circ_{\text{soln}}$, is $70.1\text{ J/(mol}\cdot\text{K)}$ at 25°C . Using the information in the table above, calculate the absolute molar entropy, S° , of aqueous urea.
 - (g) Using particle-level reasoning, explain why $\Delta S^\circ_{\text{soln}}$ is positive for the dissolution of urea in water.
 - (h) The student claims that ΔS° for the process contributes to the thermodynamic favorability of the dissolution of urea at 25°C . Use the thermodynamic information above to support the student's claim.