

27 An example of sacrificial protection is the fitting of zinc blocks to the outside of a ship's steel hull.

Which statement explains why zinc is used to protect the iron in the steel from rusting?

A Zinc is more reactive than iron so it loses electrons more easily.

B Zinc is less reactive than iron so it loses electrons more easily.

C Zinc is more reactive than iron so it gains electrons more easily.

D Zinc is less reactive than iron so it gains electrons more easily.