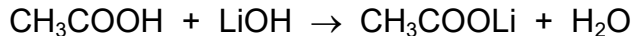


- 8 The equation for the reaction between dilute ethanoic acid and aqueous lithium hydroxide is shown.



In an experiment,  $20.0 \text{ cm}^3$  of  $0.050 \text{ mol/dm}^3$  aqueous lithium hydroxide is exactly neutralised by  $12.5 \text{ cm}^3$  of dilute ethanoic acid.

Which statement is correct?

- A** The amount of lithium hydroxide is equal to  $0.050 \times \frac{1000}{20}$  mol.
- B** The amount of ethanoic acid reacted is  $0.0010$  mol.
- C** The concentration of the dilute ethanoic acid is  $0.16 \text{ mol/dm}^3$ .
- D** The volume of aqueous lithium hydroxide used is  $0.20 \text{ dm}^3$ .