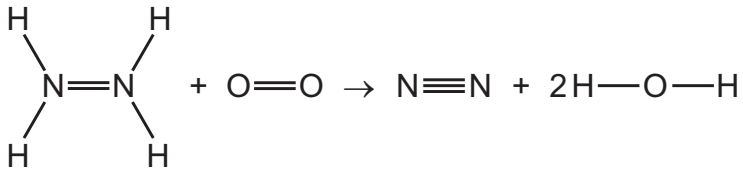
14 The equation for the combustion of hydrazine, N_2H_4, is shown.



bond	bond energy in kJ / mol
N–H	391
N=N	409
N≡N	944
O–H	463
O=O	496

What is the overall enthalpy change for the combustion of one mole of hydrazine?

A –327 kJ / mol

B –111 kJ / mol

C +111 kJ / mol

D +327 kJ / mol