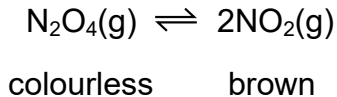


- 18 The equation shows the equilibrium between dinitrogen tetroxide,  $\text{N}_2\text{O}_4$ , and nitrogen dioxide,  $\text{NO}_2$ .

The colours of the reactant and product are shown.



The forward reaction is endothermic.

Which statement is **not** correct?

- A At equilibrium, the concentrations of the reactant and the product stay constant.
- B At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction.
- C When the pressure is increased, a darker brown colour is seen.
- D When the temperature is increased, a darker brown colour is seen.